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Re-typed from The Ultimate Chemical Equations Handbook by Hague and Smith 7. Free elements, no matter how complex the molecule, have an oxidation number (valence or charge) equal to zero. O₂, Fe, T₂, He, < /1 1 ~ ~ L-td. 8. The following are diatomic or polyatomic elements in nature which must be committed to memory.

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Note: Not all of the reactions will occur.
For those that do not, write no reaction.

1. A piece of copper is dropped into a
container of water. No reaction
2. Liquid bromine is added to a container of
sodium iodide crystals. $\text{Br (l)} + 2\text{NaI(s)}$

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$2\text{NaBr}(s) + 12(s)$ 3. An aluminum strip is immersed in a solution of silver nitrate.

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Pg. 50, Practice-Pg. 51 • Acid Base
Neutralization Reactions- Pg. 53,
Practice-Pg.54 oWeak acids and bases
do not break apart and are in net ionic
equation

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by Hague and Smith 7 Free elements, no matter how complex the molecule, have an oxidation number (valence or charge) equal to zero OJ Fe T~ He < /1 1 ~ ~ L-td
8 The following are diatomic or polyatomic

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Inc. Correlating questions (ii) from:
Rodney Schmitz (2007) Moore County
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each of the following eight reactions, in part (i) write a BALANCED equation and in part (ii) answer the question about the reaction.

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Handbook SnCO tin(II) carbonate sodium

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hydrogen carbonate NaHCO_3 1. 2. 3. 4. 5.
6. 8. 9. 10. 11. vanadium(V) oxide H_2O
dihydrogen monoxide (NH) CO
ammonium oxalate 4 224 polonium(VI)
thiocyanate $\text{Po}(\text{SCN})_6$ tetraphosphorus
decaoxide P_2O_5 10 12. 13. 15. 16. 17.
18. 19. manganese(VII) oxide copper(II)
dihydrogen phosphate

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The Ultimate Chemical Equations Handbook Determine the oxidation number of the underlined element:

KMnO_4 Since K is an alkali metal, its charge must be $1+$ Oxygen is $2-$ but there are four of them, therefore, 4 times $2-$ equals $8-$ If $1+$ and $8-$ are added Ultimate Chemical Equation

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